

ActivPhysics simulation 8.14

- Q1 A 2nd
 B 3rd
 C 1st
 D 4th

Q2 $P = nRT/V$

Q3 A: $\Delta U = 0$
 $W = Q$
 $W = \int P dV = \int \frac{nRT}{V} dV = nRT \ln V_f/V_i = 2.67 \text{ kJ}$
 $= 1.0 \cdot 8.3145 \cdot 500 \ln 19/10$

B: $Q = 0$
 $W = -\Delta U = -\frac{3}{2} nRT = 2.49 \text{ kJ}$

C: $\Delta U = 0$
 $Q = W = nRT \ln V_f/V_i = 1.0 \cdot 8.3145 \ln \frac{21}{40} = -1.61 \text{ kJ}$

D: $Q = 0$
 $W = -\Delta U = -\frac{3}{2} nRT = 2.49 \text{ kJ}$

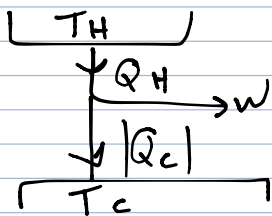
Q4 $e = \frac{W}{|Q_H|} = \frac{2.67 \text{ kJ} + 2.49 \text{ kJ} - 1.61 \text{ kJ} - 2.49 \text{ kJ}}{2.67 \text{ kJ}} = 0.397$

Q5 $e = 1 - \frac{T_c}{T_H} = 1 - \frac{300}{500} = 0.4$

Q7 $e_{\max} = 1 - \frac{T_{\text{coldest}}}{T_{\text{hottest}}} = 1 - \frac{270 \text{ K}}{600 \text{ K}} = 0.55$

$e_{\min} = 1 - \frac{T_{\text{warmest cold}}}{T_{\text{coolest hot}}} = 1 - \frac{350 \text{ K}}{400 \text{ K}} = 0.13$

Heat engine

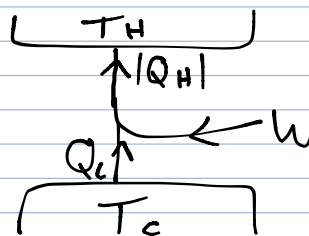


$W = Q_H + Q_C$
 $= |Q_H| - |Q_C|$

$Q_H > 0$ $Q_C < 0$
 $W > 0$

efficiency = $\frac{\text{output work}}{\text{input heat}} = \frac{W}{Q_H}$

Refrigerator



$W = Q_H + Q_C$
 $|W| = |Q_H| - |Q_C|$

$Q_H < 0$ $Q_C > 0$
 $W < 0$

coeff. of performance = $\frac{\text{heat removed}}{\text{work required}} = \frac{|Q_C|}{|W|} = K$

example A freezer has $K=2.40$ and brings 1.80 kg of H_2O at 25.0°C down to -5.0°C in 60 min .

a) How much heat is removed?

$$\begin{aligned} Q_c &= Q_{\text{cool H}_2\text{O}} + Q_{\text{convert to ice}} + Q_{\text{cool ice}} \\ &= mc\Delta T_{\text{H}_2\text{O}} - mL_f + mc\Delta T_{\text{ice}} \\ &= 1.80\text{ kg} \left(\frac{4190\text{ J}}{\text{kg}\cdot\text{K}} (-25\text{ K}) - 334 \cdot 10^3 \frac{\text{J}}{\text{kg}} + 2010 \frac{\text{J}}{\text{kg}\cdot\text{K}} (-5\text{ K}) \right) = -8.08 \cdot 10^5 \text{ J} \end{aligned}$$

b) How much electrical energy is consumed?

$$W = \frac{|Q_c|}{K} = \frac{8.08 \cdot 10^5 \text{ J}}{2.40} = 3.37 \cdot 10^5 \text{ J}$$

c) How much heat is delivered to the room?

$$|Q_H| = |W| + |Q_c| = 3.37 \cdot 10^5 + 8.08 \cdot 10^5 \text{ J} = 1.14 \cdot 10^6 \text{ J}$$

ch 04 s2.html if $|Q_c|$ were 0, then $W = Q_H$ and $e = 1.0$ ← not possible
 \Rightarrow "A" works

if Q_H were smaller than W , then $e > 1.0$

\Rightarrow "B" works

$\rightarrow E: A \nexists B$

53.html

refrigerator perspective: no work is needed to cool the house

entropy perspective: going from disordered to more ordered \rightarrow not allowed

Polleverywher: do the exam prep / review on

A: Friday
B: Wednesday