Heat Problem

1. 50 mL of ethanol at -18 °C is mixed with 300 mL of water at 20 °C. Find the temperature of the mixture once it has reached thermal equilibrium. The density of water and ethanol are 1 g/mL and 0.789 g/mL, respectively, and their specific heats are 4186 J/kg K and 2450 J/kg K, respectively.

Latent Heat Problem

2. Calculate the heat to convert 5 kg of ice at -20 °C into steam at 100 ° C.

		L_f			L_v	
Substance	**Melting point (°C)**	**kJ/kg**	**kcal/kg**	**Boiling point (°C)**	**kJ/kg**	**kcal/kg**
Helium	−269.7	5.23	1.25	−268.9	20.9	4.99
Hydrogen	−259.3	58.6	14.0	−252.9	452	108
Nitrogen	−210.0	25.5	6.09	−195.8	201	48.0
Oxygen	−218.8	13.8	3.30	−183.0	213	50.9
Ethanol	−114	104	24.9	78.3	854	204
Ammonia	−75		108	−33.4	1370	327
Mercury	−38.9	11.8	2.82	357	272	65.0
Water	0.00	334	79.8	100.0	2256 [4]	540[5]

Substances	**Specific heat (c)**	
Solids	J/(kg·°C)	kcal/(kg°C)[1]
Aluminum	900	0.215
Asbestos	800	0.19
Concrete, granite (average)	840	0.20
Copper	387	0.0924
Glass	840	0.20
Gold	129	0.0308
Human body (average at 37 °C)	3500	0.83
Ice (average, -50°C to 0°C)	2090	0.50
Iron, steel	452	0.108
Lead	128	0.0305
Silver	235	0.0562
Wood	1700	0.4
Liquids		
Benzene	1740	0.415
Ethanol	2450	0.586
Glycerin	2410	0.576
Mercury	139	0.0333
Water (15.0 °C)	4186	1.000